

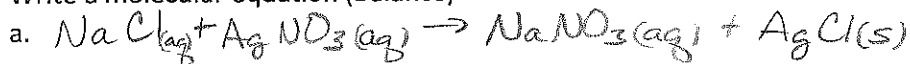
Standard: (#4-2)

Solubility Reactions Draw

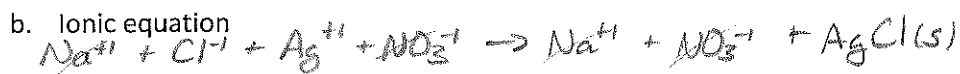
1. Equal concentrations of NaCl reacts with AgNO₃ causing a white solid to appear.

Spectators: Na⁺ NO₃⁻

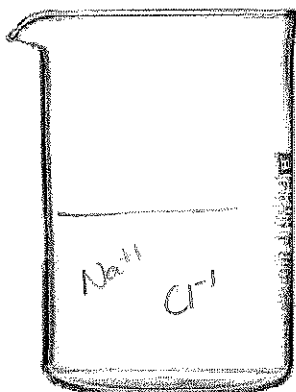
Write a molecular equation (Balance)



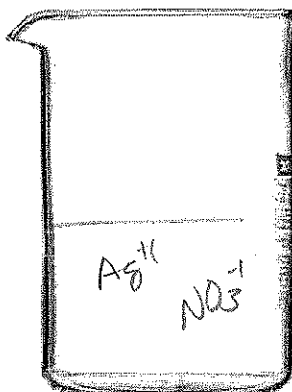
b. Ionic equation



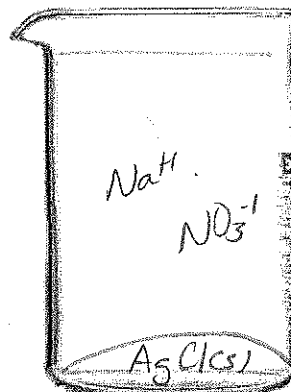
Net ionic equation (Balance)



d. NaCl



AgNO₃



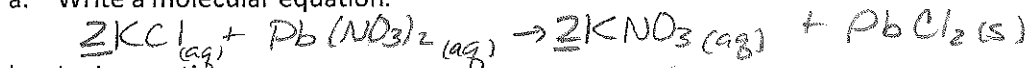
Products

2. Equal concentrations of KCl reacts with Pb(NO₃)₂ causing a white solid to appear.

Spectators: K⁺ + NO₃⁻

(Balance all equations)

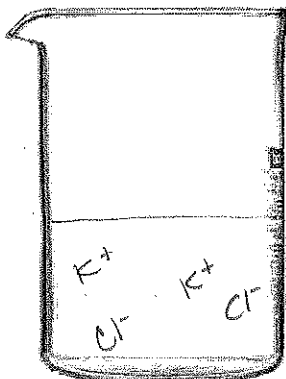
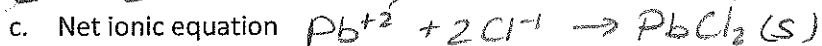
a. Write a molecular equation.



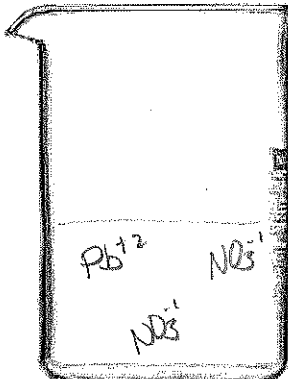
b. Ionic equation



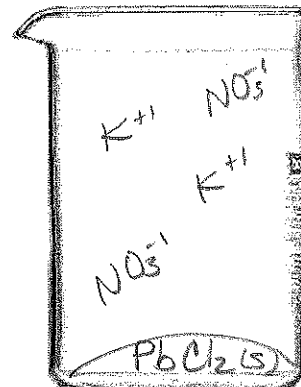
c. Net ionic equation



e. KCl



Pb(NO₃)₂



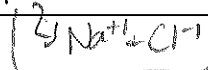
Products

3. Zinc Chloride is mixed with a solution of sodium Carbonate that is ^{same} twice as concentrated ~~on~~
 (Balance all equations)

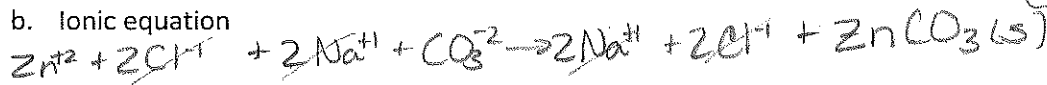
a. Write a molecular equation.



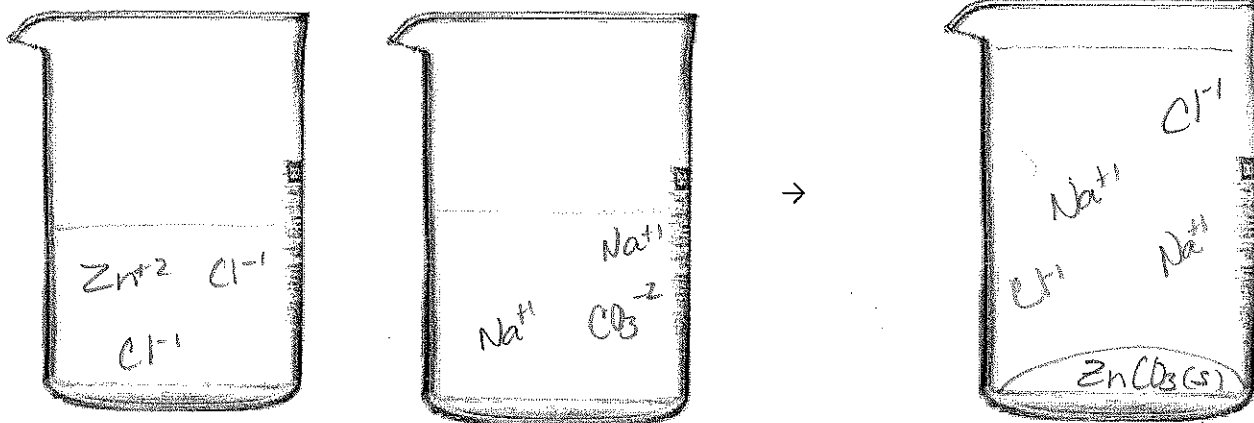
Spectators:



b. Ionic equation



c. Net ionic equation $\text{Zn}^{+2} + \text{CO}_3^{-2} \rightarrow \text{ZnCO}_3(\text{s})$



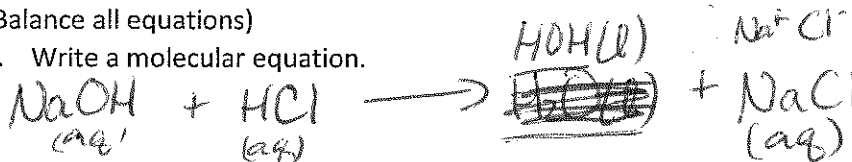
d. ~~NaCl~~

~~AgNO₃~~

Products

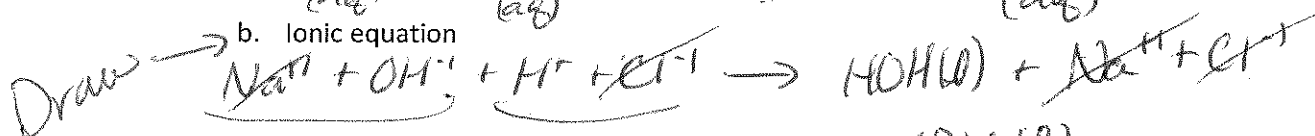
4. Sodium Hydroxide is mixed with a solution of Hydrochloric acid that is equally concentrated.
 (Balance all equations)

a. Write a molecular equation.



Spectators: ~~Na⁺ Cl⁻~~

b. Ionic equation



c. Net ionic equation

